



Science 1206 (12-13)

Quiz 2 Naming/Formula Writing

Student Name: _____

1. For the following substances:
- In the first column, class the compound as ionic (I) or molecular (M).
 - Give the correct IUPAC name in the space provided; use common names as necessary.

PRINT ALL ANSWERS CLEARLY. Any unclear numbers or letters will cause a response to be marked incorrect.

I or M	Formula	Name
1 M	N_2O_5	dinitrogen pentoxide
2 M	Br_2	bromine
3 I	K_2CO_3	potassium carbonate
4 I	FeSO_4	iron (II) sulfate
5 I	MgBr_2	magnesium bromide
6 I	$\text{Zn}(\text{MnO}_4)_2$	zinc permanganate
7 I	CoHSO_4	cobalt hydrogen sulfate
8 I	$\text{Al}_2(\text{CrO}_4)_3$	Aluminium
9 M	S_2F	disulfur monofluoride
10 M	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$	sucrose
11 M	HCl_4	hydrogen tetrachloride
12 M	PN_8	phosphorous octanitride
13 I	Li_2S	lithium sulfide
14 I	$(\text{NH}_4)_2\text{SO}_4$	ammonium sulfate
15 M	NH_3	ammonia
16 I	Ag_2O	silver oxide
17 M	H_2	hydrogen
18 M	NO_2	nitrogen dioxide
19 I	MnO_2	manganese (IV) oxide
20 I	CaCl_2	calcium chloride

2. For the following substances:
- in the first column, class the compound as ionic (I) or molecular (M).
 - give the correct chemical *formula* in the space provided; some formulas correspond to traditional names.

PRINT ALL ANSWERS CLEARLY. Any unclear numbers or letters will cause a response to be marked incorrect.

I or M	Name	Formula
1 I	magnesium oxalate	MgC_2O_4
2 I	potassium oxide	K_2O
3 I	iron(III) sulfide	Fe_2S_3
4 M	nitrogen trifluoride	NF_3
5 M	ozone	O_3
6 I	magnesium hydrogen carbonate	$\text{Mg}(\text{HCO}_3)_2$
7 I	sodium borate	Na_3BO_3
8 M	fluorine	F_2
9 I	carbon dioxide	CO_2
10 I	ammonium sulfide	$(\text{NH}_4)_2\text{S}$
11 I	lithium dichromate	$\text{Li}_2\text{Cr}_2\text{O}_7$
12 I	tin(IV) sulfide	SnS_2
13 M	Diphosphorus pentoxide	P_2O_5
14 I	zinc thiosulfate	ZnS_2O_3
15 I	aluminum permanganate	$\text{Al}(\text{MnO}_4)_3$
16 I	nickel(III) sulfide	Ni_2S_3
17 I	dinitrogen tetrahydride	N_2H_4
18 I	sodium benzoate	$\text{NaC}_6\text{H}_5\text{COO}$
19 I	barium cyanide	$\text{Ba}(\text{CN})_2$
20 M	silicon tetrafluoride	SiF_4